# Superparamagnetic properties of γ-Fe<sub>2</sub>O<sub>3</sub> particles: Mössbauer spectroscopy and DC magnetic measurements

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### **Synopsis**

Small particles and transition to superparamagnetism

Magnetic measurements and superparamagnetism - blocking temperatures T<sub>B</sub>

Nanoparticles of Fe<sub>2</sub>O<sub>3</sub> oxide

Synthesis and characterization

Mössbauer spectroscopy and magnetic properties

Mössbauer spectra at various temperatures

Analysis of ZFC/FC magnetic moments

Distribution of blocking temperatures



Comparison with Mössbauer spectra

### **Conclusions**

#### Small particles and transition to superparamagnetism



Each <sup>57</sup>Fe atom with given SRO

Time average of **B** over  $10^{-7}s \rightarrow 0$ Sextet(s) broaden(s) and finally collapse to doublet/singlet



Spontaneous change of direction of



For **superparamagnetism** decisive K.V

Anisotropy constant K and particle volume V



atomic moments : paramagnetism

particle moments : superparamagnetism

vs **k**<sub>B</sub>.**T** (Boltzmann constant x Temperature)

Note different characteristic time in Mössbauer Spectroscopy (~10<sup>-7</sup>s for <sup>57</sup>Fe)

and

D.C. magnetic measurements (seconds)



## Blocking temperatures and their distribution



## Experimental

#### Synthesis of dextran-stabilized $\gamma$ -Fe<sub>2</sub>O<sub>3</sub> (maghemite?) nanoparticles

Magnetic **iron oxide colloid** was prepared by **chemical coprecipitation** of Fe(II) and Fe(III) salts with ammonium hydroxide in (Carboxymethyl)-dextran solution and subsequent **oxidation** with sodium **hypochlorite** to **avoid uncontrolled oxidation** of primarily formed magnetite **by air oxygen**.

**Precipitation** method produces very fine **nanoparticles**, with **CM-dextran** providing **sterical stabilization** of magnetic nanoparticles: COOH group possesses affinity to Fe3+ ions, the dextran part allows the dispersion in water and **prevents aggregation** of the particles forming a **stable colloid**.

Colloidal **magnetic phase** precipitated in the presence of (carboxymethyl)dextran **formed "chains"** of fine **particles about 5 nm** in **diameter** and with **narrow** size **distribution** (PDI = 1.06). (Fig. 1)

Magnetic poly(glycidyl methacrylate)microspheres containing maghemite prepared by emulsion polymerization



## Experimental cntnd

Phase compositions of the magnetic iron oxide particles (and PGMA composite microspheres) and

**size** of the magnetic crystallites determined by **X-ray powder diffraction** (Bruker D8 diffractometer equipped with CuK, Sol-X energy dispersive detector). XRD patterns analyzed with the Rietveld method using the FULLPROF program.

Instrumental, strain and size **contributions** to peak **broadening** resolved using Thompson-Cox-Hastings pseudo-Voigt profile. Instrumental resolution was **checked** by measuring **tungsten powder** with grain size of **9.4 µm**.

**Magnetic properties** were measured by **SQUID** magnetometer (MPMS5 by Quantum Design, USA) in the temperature range 5-295 K in the fields up to **5** T.

The Mössbauer spectra acquired in the transmission mode with 57Co/Rh source moving with constant acceleration. Calibration by a standard  $\alpha$ –Fe foil and I.S. expressed with respect to this standard at 293 K.

Samples were measured in a Janis cryosystem at temperatures 120 – 300K. The spectra were fitted with the help of the NORMOS program.

#### Moessbauer spectra at room temperature and 120 K



#### Parameters of the Mössbauer spectra

	RT			120 K		
	B <sub>hf</sub>	ISO	r. a. %	B <sub>hf</sub>	ISO	r.a. %
Sextet 1	40.3	0.38	14.4	47.9	0.13	38.3
Sextet 2				44.0	0.10	41.8
Broad sextet	25.9	0.29	76.0	26.4	0.34	16.7
Doublet 1		0.37	5.0		~0.0	3.2
Doublet 2		0.18	4.6			
mostly ordered						

mostly superparamagnetic

maghemite

phase? Fe<sup>3+</sup> !

#### Temperature dependence of Mössbauer spectra



		Sextet 1	Sextet 2	Broad Sextet	Doublet 1	Doublet 2
Room Temp.	B <sub>hf</sub> [T]		40.3	25.9		
	ISO [mm/s]		0.38	0.29	0.37	0.18
	r.a. [%]		14.4	76.0	5.0	4.6
275 K	B <sub>hf</sub> [T]		40.4	28.9		
	ISO [mm/s]		0.31	0.29	0.41	0.26
	r.a. [%]		25.3	44.9	22.1	7.7
230 K	B <sub>hf</sub> [T]		42.4	30.3		
	ISO [mm/s]		0.39	0.46	0.51	0.19
	r.a. [%]		40.6	38.1	15.2	6.1
200 K	B <sub>hf</sub> [T]	46.3	43.1	30.2		
	ISO [mm/s]	0.34	0.37	-0.41	0.75	-0.4
	r.a. [%]	3.2	44.7	36.7	10.4	5.0
170 K	B <sub>hf</sub> [T]	46.2	42.0	29.7		
	ISO [mm/s]	0.39	0.51	0.2	0.11	0.31
	r.a. [%]	26.7	29.9	31.2	10.5	1.7
140 K	B <sub>hf</sub> [T]	48.3	44.7	30.5		
	ISO [mm/s]	0.45	0.4	0.06		-0.03
	r.a. [%]	37.7	40.4	19.1		2.8
120 K	B <sub>hf</sub> [T]	47.9	44.0	26.4		
	ISO [mm/s]	0.13	0.1	0.34		-0.03
	r.a. [%]	38.3	41.8	16.7		3.2







#### **Relaxation time and Energy barrier**

	$\tau = \tau_0 \exp(\Delta E / kT)$		with $\Delta E = K.V$		for uniaxial magnetocrystalline (spherical particles) anisotropy			
Inni	ıt data:							
πραι σαιά.			MS:		$ au \approx 10^{-7}s$ , T $\Box$ (200, 250 K)			
			ZFC/F	C curves:	$ au \approx 1s$	T 🗆 (*	I3, 70 K)	
V: d ≈ 5 nm cube: 1.25 x 10 <sup>-25</sup> m <sup>3</sup>								
				sphere	e: 6.5 x ′	10 <sup>-26</sup> m <sup>3</sup>		
Assumption: K independent of temperature								
		ę	gives $\tau_0$	≈ 4.2 x 10	<sup>-8</sup> s - 1.7	∕ x 10 <sup>-11</sup> s		
а	nd <mark>K</mark> ≈	2.4 x <sup>-</sup>	<b>10</b> <sup>4</sup> - 3.	6 x 10 <sup>5</sup> J/r	m <sup>3</sup> (=	) 2.4 x 10	<sup>5</sup> – 3.6 x 10	) <sup>6</sup> erg/cm <sup>3</sup>

which plausibly compares in absolute value with LB data

 $K_1 = -2.5 \times 10^5 \text{ erg/cm}^3$  (for cubic crystal)



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### TEM pictures of PGMA microspheres with various concentrations of Fe<sub>2</sub>O<sub>3</sub>



0.8 wt % of Fe<sub>2</sub>O<sub>3</sub>

6.2 wt % of Fe<sub>2</sub>O<sub>3</sub>